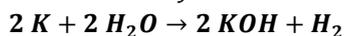


NAME: _____ DATE: _____ PERIOD: _____

- d. Determine the number of moles of hydrogen produced when 0.0400 mole of potassium is used. *Use dimensional analysis and include all units!*



- e. **Stoichiometry is a _____ to _____ relationship.**

4. Sodium chloride or table salt (*NaCl*) is produced when chlorine and sodium react vigorously together.

- a. Write the balanced equation for the reaction above:

- b. Write the mole ratio between *NaCl* and *Cl₂* in the form of a fraction from the balanced equation above:

- c. If 1.25 moles of chlorine gas react in the equation above, what would be the mass of *NaCl* produced? *Use dimensional analysis and include all units!*

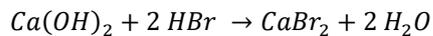
Step 1 – Convert moles of *Cl₂* to moles of *NaCl*

Step 2 – Convert moles of *NaCl* into grams

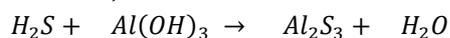
NAME: _____ DATE: _____ PERIOD: _____

CHECKPOINTS

1. How many moles of water are produced when 3.21 moles of calcium hydroxide are used in the reaction below?



2. How many moles of aluminum sulfide are formed when you react 0.45 moles of hydrogen sulfide? (*Balance the equation first*)



3. How many moles of water and sodium nitride are formed when you react 0.75 mol of sodium hydroxide with ammonia (hydrogen nitride).